	Class A	Class B	Class C	Class D
Date Opened	3 Months	6 Months	6 Months	Only if peroxide crystals are present.
Date Received	1 Year	1 Year	1 Year	

The list below gives examples of peroxide-forming materials by Class. Materials listed are samples of Class A, B, C and D peroxide-forming materials, but are not all-inclusive.

Spontaneously Decpmpose and become explosive with exposure to air without concentration.

- Butadiene (liquid monomer)
- Isopropyl ether
- Sodium amide (sodamide)
- Chloroprene (liquid monomer)
- Potassium amide
- Tetrafluoroethylene (liquid monomer)
- Divinyl acetylene
- Potassium metal
- Vinylidene chloride

Require external energy for spontaneous decomposition. Form explosive peroxides when distilled, evaporated or otherwise concentrated.

- Acetal
- Acetaldehyde
- Acrylonitrile
- Benzyl alcohol
- Cumene
- Cyclohexanol
- Chlorotrifluoroethylene
- Chloroprene
- Cyclohexene
- Decahydronaphthalene
- Diacetylene
- Dicyclopentadiene
- Diethyl ether
- Diethylene glycol dimethyl ether (diglyme)
- Dioxanes
- Ethylene glycol dimethyl ether (glyme)
- Furan
- Methyl isobutyl ketone
- Methylacetylene
- Methylcyclopentane
- Methyl methacrylate
- Other secondary alcohols
- Tetrahydrofuran
- Tetrahydronaphthalene
- Vinyl ethers
- 1-Phenylethanol
- 2-Butanol
- 2-Cyclohexen-1-ol
- 2-Hexanol
- 2-Pentanol
- 2-Phenylethanol
- 3-Methyl-1-butanol
- 4-Heptanol
- 4-Methyl-2-pentanol
- 4-Penten-1-ol

Highly reactive and can auto-polymerize as a result of internal peroxide accumulation. The peroxides formed in these reactions are extremely shock and heat sensitive.

- Acrylic acid
- Chlorotrifluoroethylene
- Vinyl acetate
- Acrylonitrile
- Methyl methacrylate
- Vinylacetylene (gas)